



Cambridge Assessment International Education
Cambridge International General Certificate of Secondary Education

COMBINED SCIENCE

0653/21

Paper 2 Multiple Choice (Extended)

October/November 2019

45 minutes

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)



READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

Electronic calculators may be used.

This document consists of **15** printed pages and **1** blank page.

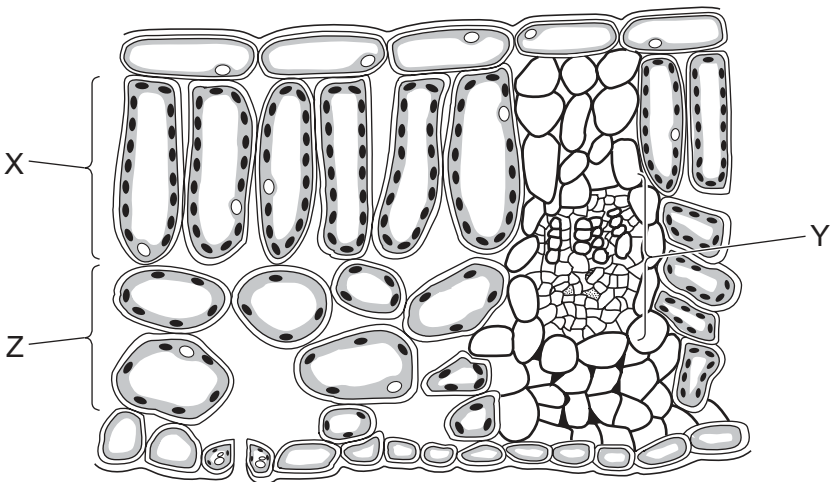


1 A biologist keeps a potted plant in a laboratory.

Which feature of the potted plant shows that it is a living organism?

- A It grows larger over time.
- B It has green leaves.
- C The compost in the pot dries after he waters it.
- D The stems contain xylem.

2 The diagram shows a section through a leaf.



Which row correctly identifies the labelled parts of the leaf section?

	X	Y	Z
A	cuticle	vascular bundle	palisade mesophyll
B	palisade mesophyll	vascular bundle	spongy mesophyll
C	palisade mesophyll	cuticle	spongy mesophyll
D	spongy mesophyll	cuticle	vascular bundle

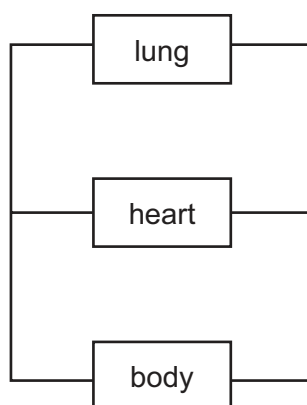
- 3 1 cm³ of substance **X** is added to 10 cm³ starch suspension and mixed. Food tests are carried out immediately after mixing and again after an hour.

The results of the tests are shown in the table.

test reagent	colour of solution after mixing	colour of solution after one hour
Benedict's solution	blue	orange
iodine solution	blue / black	brown

What is substance **X**?

- A** amylase
B protease
C lipase
D sugar
- 4 How are root hair cells adapted for absorption of water?
- A** large surface area
B thick cell wall
C many chloroplasts
D no nucleus
- 5 The diagram shows a double circulatory system.



In which direction does the blood flow in this type of system?

- A** heart → body → heart → lung
B body → heart → body → lung
C heart → body → lung → heart
D lung → heart → lung → body

- 6 The concentrations of carbon dioxide and oxygen in expired air differ from the concentrations in inspired air.

	gas	concentration in expired air
1	carbon dioxide	higher
2	carbon dioxide	lower
3	oxygen	higher
4	oxygen	lower

Which rows correctly show the difference?

- A** 1 and 3 **B** 1 and 4 **C** 2 and 3 **D** 2 and 4
- 7 Which statement about aerobic respiration is correct?
- A** It exchanges gases through the walls of the alveoli.
B It expels carbon dioxide from the lungs.
C It only produces carbon dioxide and energy.
D It uses oxygen to release energy from glucose.

- 8 What is the effect of adrenaline on the rate of breathing and pulse rate?

	rate of breathing	pulse rate
A	decreases	decreases
B	decreases	increases
C	increases	decreases
D	increases	increases

- 9 Auxins are produced in the tip of a shoot.

In an experiment, auxin was applied to a plant shoot just below its tip.

Which row describes the change to the shoot and explains this change?

	shoot	explanation
A	becomes longer	cells divide
B	becomes longer	cells elongate
C	becomes longer	cells elongate and then divide
D	did not change length	the auxin has no effect

10 Which statement about sexual reproduction is **always** correct?

- A It involves only one parent.
- B It involves the fusion of nuclei.
- C It produces genetically identical offspring.
- D It takes place only in animals.

11 The table gives comparisons between insect-pollinated and wind-pollinated flowers.

Which row is correct?

	surface of stigma		surface of pollen	
	insect	wind	insect	wind
A	feathery	sticky	smooth	spiky
B	feathery	sticky	spiky	smooth
C	sticky	feathery	smooth	spiky
D	sticky	feathery	spiky	smooth

12 The following are adaptive features of some gametes.

- 1 a flagellum
- 2 a jelly coat
- 3 a sac of enzymes at one end
- 4 a store of energy

Which are features of a human male gamete?

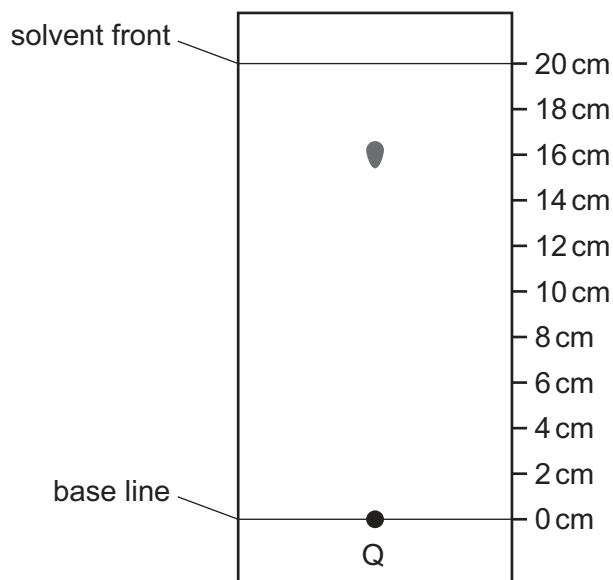
- A** 1 and 3 **B** 1 and 4 **C** 2 and 3 **D** 2 and 4

13 An increased volume of nitrate ions is washed into a lake.

Why does this result in a decrease in the number of fish in the lake?

- A There is a decrease in the decomposition of producers.
- B There is a decrease in the growth of producers.
- C There is an increase in aerobic respiration by decomposers.
- D There is an increase in dissolved oxygen.

14 A chromatogram of substance Q is shown.



What is the R_f value of Q?

- A** 0.2 **B** 0.4 **C** 0.8 **D** 1.6

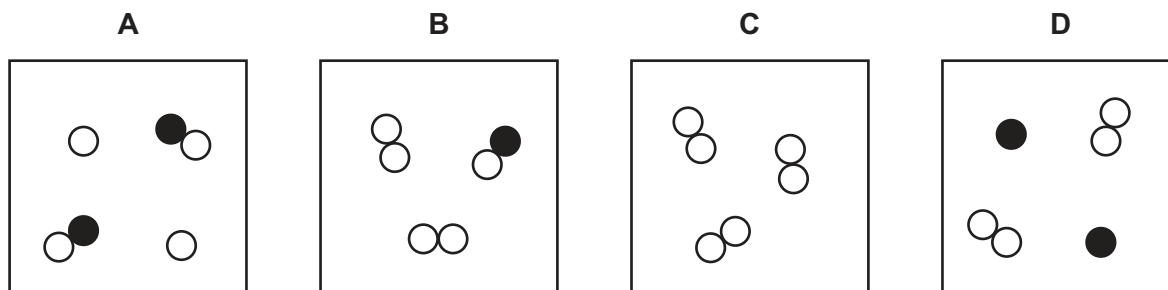
15 Four processes are listed.

- 1 melting of ice
- 2 electrolysis of molten lead(II) bromide
- 3 combustion of carbon
- 4 rusting of iron

Which processes are chemical changes?

- A** 1 and 3 only **B** 1, 2 and 3 **C** 2 and 4 only **D** 2, 3 and 4

16 Which diagram represents a mixture of two different elements?



17 The fertiliser ammonium sulfate has the formula $(\text{NH}_4)_2\text{SO}_4$.

How many atoms of each element are present in the formula?

	number of hydrogen atoms	number of nitrogen atoms	number of oxygen atoms	number of sulfur atoms
A	4	1	1	1
B	4	2	4	1
C	8	1	4	1
D	8	2	4	1

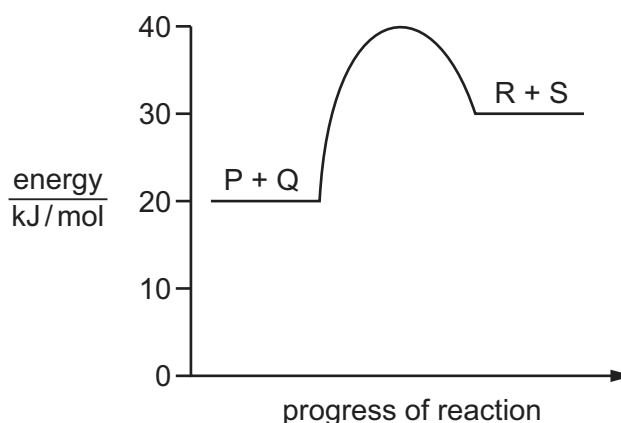
18 Element X is a non-metal used in the treatment of the water supply.

It is made during the electrolysis of a metal salt.

What is the colour of X and at which electrode is it made?

	colour	electrode
A	red	anode
B	red	cathode
C	yellow-green	anode
D	yellow-green	cathode

19 The energy level diagram for the reaction between P and Q to produce R and S is shown.



Which statement about this reaction is correct?

- A** Energy is required to form new bonds.
- B** The activation energy is 10 kJ/mol.
- C** The energy required to break bonds is greater than the energy given out when bonds form.
- D** The reaction is exothermic because the energy of the products is greater than the energy of the reactants.

- 20 Calcium carbonate reacts with dilute hydrochloric acid.

The time taken to collect 10 cm³ of carbon dioxide is recorded.

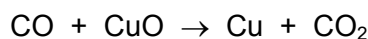
The experiment is repeated at a different temperature. The results are shown.

experiment	temperature /°C	time taken /s
1	20	55
2	80	30

The rate of reaction in each experiment is different.

Which statement about the rate of reaction of experiment 1, compared with experiment 2, is correct?

- A It is greater because at the lower temperature the particles move more slowly so they have more time to react.
 - B It is greater because the particles collide more frequently.
 - C It is lower because the particles collide at the same frequency and fewer of them have the minimum energy to react.
 - D It is lower because the particles collide less frequently and fewer of them have the minimum energy to react.
- 21 The equation for the reaction of carbon monoxide with copper oxide is shown.

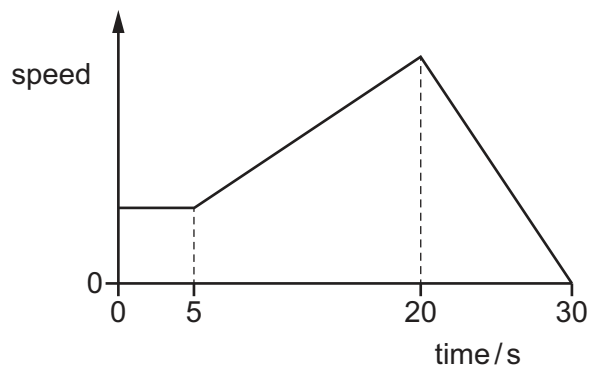


Which statement about this reaction is **not** correct?

- A Carbon dioxide is the oxidising agent.
 - B Carbon monoxide is the reducing agent.
 - C Carbon monoxide is being oxidised.
 - D Copper oxide is the oxidising agent.
- 22 Which two substances are used to make copper sulfate?
- A copper and dilute sulfuric acid
 - B copper and sulfur
 - C copper oxide and dilute sulfuric acid
 - D copper oxide and sulfur

- 23 Which statement about elements in the Periodic Table is **not** correct?
- A Elements in Group I react by gaining electrons to form ions.
 - B Elements in Group II have two electrons in their outer shell.
 - C Elements in Group VIII are unreactive because they have full outer shells of electrons.
 - D Elements with seven outer shell electrons are non-metals.
- 24 Which statement about alloys is correct?
- A They are made from metals because metals are poor electrical conductors.
 - B They are mixtures of compounds that contain metals.
 - C They have all the same properties as the metals from which they are made.
 - D They have different properties to the metals from which they are made.
- 25 Which gas is **not** present in clean air?
- A carbon monoxide
 - B neon
 - C nitrogen
 - D water vapour
- 26 Which two gases cause an enhanced greenhouse effect when their concentrations in the atmosphere increase?
- A carbon monoxide and carbon dioxide
 - B carbon dioxide and methane
 - C methane and sulfur dioxide
 - D sulfur dioxide and carbon monoxide
- 27 Which statement about fractions obtained from petroleum is **not** correct?
- A Different fractions have different boiling points.
 - B Gasoline contains molecules of only one hydrocarbon.
 - C Molecules in diesel oil are smaller than molecules in bitumen.
 - D Naphtha is used as a feedstock for making chemicals.

- 28 The graph shows how the speed of a car changes with time. The car travels at constant speed, then accelerates, and finally brakes to a stop.



The car travels 60 m while it brakes to a stop.

What is the average speed of the car while it is braking?

- A** 3.0 m/s **B** 4.0 m/s **C** 6.0 m/s **D** 12 m/s

- 29 The gravitational field strength on Mars is less than that on Earth.

An object is taken from Earth to Mars.

Which statement describes the object when it is on Mars?

- A** It has greater mass than on Earth.
B It has less mass than on Earth.
C It has less weight than on Earth.
D It has the same weight as on Earth.

- 30** A student wants to determine the density of an irregularly shaped stone. He has a measuring cylinder, water, a balance and a thermometer.

These are his measurements:

- P initial volume of water in measuring cylinder
- Q mass of measuring cylinder
- R mass of stone
- S mass of water in measuring cylinder
- T reading on measuring cylinder with stone fully immersed in the water
- U temperature of water

Which three measurements are needed to determine the density of the stone?

- A** P, R and T
 - B** P, T and U
 - C** Q, R and S
 - D** Q, S and U
- 31** Which device uses a non-renewable energy source?

- A** diesel engine
- B** solar cell
- C** water turbine
- D** windmill

- 32** A metal pan containing water is heated on a hot stove. Energy is transferred thermally from the stove to the water.

How is the energy transferred through the pan and then throughout the water?

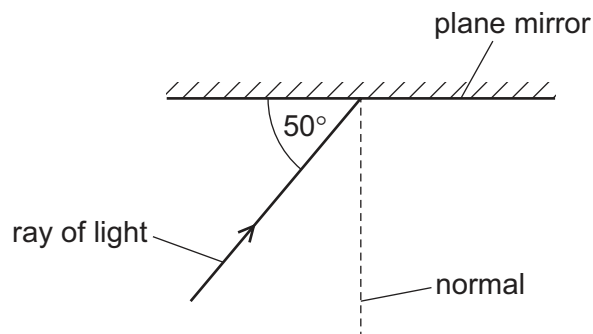
	through the pan	throughout the water
A	conduction	conduction
B	conduction	convection
C	convection	conduction
D	convection	convection

- 33** On a hot, sunny day a boy finds that his head stays cooler when he wears a white hat than when he wears an otherwise identical black hat.

Why does the white hat keep his head cooler?

- A** It absorbs less radiation from the Sun than the black hat.
 - B** It conducts less heat energy than the black hat.
 - C** It conducts more heat energy than the black hat.
 - D** It emits more radiation from his head than the black hat.
- 34** Which wave is longitudinal?
- A** infrared
 - B** radio
 - C** sound
 - D** ultraviolet

- 35** The diagram shows light striking a plane mirror.

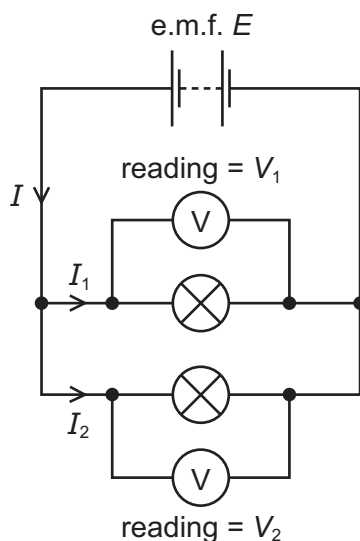


What is the angle of reflection of the ray when it is reflected from the mirror?

- A** 40°
- B** 50°
- C** 80°
- D** 100°

- 36** A circuit contains a battery of e.m.f. E , two lamps and two voltmeters, connected as shown.

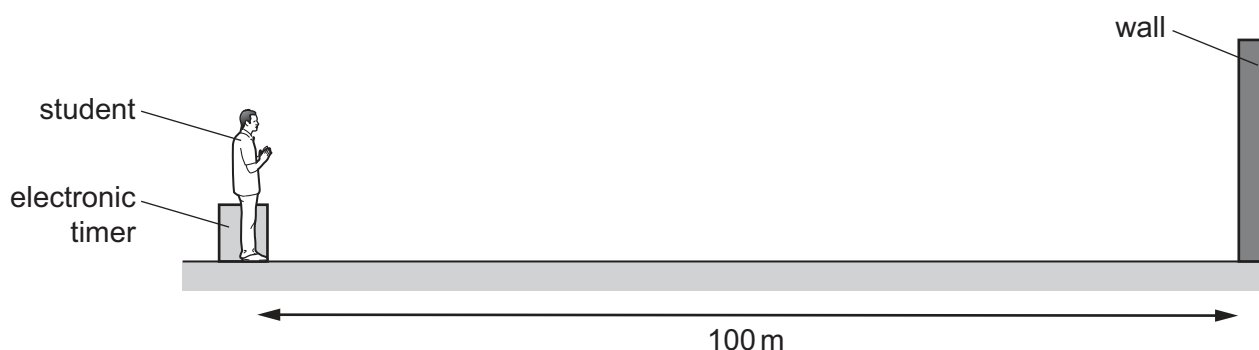
The voltmeter readings V_1 and V_2 , and the current in three parts of the circuit I , I_1 and I_2 are labelled.



Which row gives expressions for current I and e.m.f. E ?

	current I	e.m.f. E
A	$I = I_1 = I_2$	$E = V_1 = V_2$
B	$I = I_1 = I_2$	$E = V_1 + V_2$
C	$I = I_1 + I_2$	$E = V_1 = V_2$
D	$I = I_1 + I_2$	$E = V_1 + V_2$

- 37** A student measures the speed of sound. He claps his hands and the sound reflects from a wall that is 100 m away from him.



An electronic timer next to the student detects the echo of the sound 0.60 s after it is made.

Which calculation gives the speed of sound?

- A** $\frac{200}{0.30}$ m/s **B** $\frac{200}{0.60}$ m/s **C** $\frac{100}{0.60}$ m/s **D** $\frac{100}{1.2}$ m/s

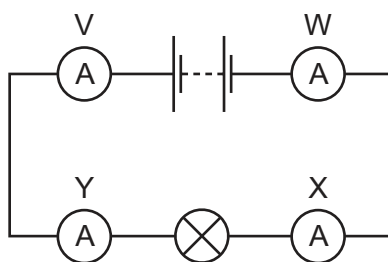
- 38** A piece of wire has a resistance of $8.0\ \Omega$.

The length of the wire is doubled and the diameter of the wire is halved.

What is the new resistance of the wire?

- A** $2.0\ \Omega$ **B** $4.0\ \Omega$ **C** $8.0\ \Omega$ **D** $64\ \Omega$

- 39** Four ammeters V, W, X and Y are connected in the circuit shown.



Which ammeters have the same reading as each other?

- A** V and W only
B V and Y only
C X and Y only
D V, W, X and Y
- 40** A 20 V power supply provides a current of 5.0 A for 1.0 minute.

How much energy does the power supply deliver?

- A** 4.0 J **B** 100 J **C** 240 J **D** 6000 J

Answer Sheet_____

	A	B	C	D
1	☐	☐	☐	☐
2	☐	☐	☐	☐
3	☐	☐	☐	☐
4	☐	☐	☐	☐
5	☐	☐	☐	☐
6	☐	☐	☐	☐
7	☐	☐	☐	☐
8	☐	☐	☐	☐
9	☐	☐	☐	☐
10	☐	☐	☐	☐
11	☐	☐	☐	☐
12	☐	☐	☐	☐
13	☐	☐	☐	☐
14	☐	☐	☐	☐
15	☐	☐	☐	☐
16	☐	☐	☐	☐
17	☐	☐	☐	☐
18	☐	☐	☐	☐
19	☐	☐	☐	☐
20	☐	☐	☐	☐

	A	B	C	D
21	☐	☐	☐	☐
22	☐	☐	☐	☐
23	☐	☐	☐	☐
24	☐	☐	☐	☐
25	☐	☐	☐	☐
26	☐	☐	☐	☐
27	☐	☐	☐	☐
28	☐	☐	☐	☐
29	☐	☐	☐	☐
30	☐	☐	☐	☐
31	☐	☐	☐	☐
32	☐	☐	☐	☐
33	☐	☐	☐	☐
34	☐	☐	☐	☐
35	☐	☐	☐	☐
36	☐	☐	☐	☐
37	☐	☐	☐	☐
38	☐	☐	☐	☐
39	☐	☐	☐	☐
40	☐	☐	☐	☐

The Periodic Table of Elements

Group																			
I	II											III	IV	V	VI	VII	VIII		
		1 H hydrogen 1																	
		Key atomic number atomic symbol name relative atomic mass																	
3 Li lithium 7	4 Be beryllium 9													5 B boron 11	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19	10 Ne neon 20
11 Na sodium 23	12 Mg magnesium 24													13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40
19 K potassium 39	20 Ca calcium 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52	25 Mn manganese 55	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59	29 Cu copper 64	30 Zn zinc 65	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84		
37 Rb rubidium 85	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium —	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131		
55 Cs caesium 133	56 Ba barium 137	57–71 lanthanoids		72 Hf hafnium 178	73 Ta tantalum 181	74 W tungsten 184	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	79 Au gold 197	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium —	85 At astatine —	86 Rn radon —		
87 Fr francium —	88 Ra radium —	89–103 actinoids		104 Rf rutherfordium —	105 Db dubnium —	106 Sg seaborgium —	108 Hs hassium —	109 Mt meitnerium —	110 Ds darmstadtium —	111 Rg roentgenium —	112 Cn copernicium —	114 Fl flerovium —		116 Lv livermorium —		—			

lanthanoids	57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium —	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
	89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).